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gas constant and the ratio of specific heats of oxygen. The units are: pressure N/m<sup>2</sup>, temperature K, area m<sup>2</sup> and mass flow rate kg/s. Assuming that the temperature of the oxygen in the bottle does not change with time, determine the time it takes to reduce the pressure to one half of its initial value.  $V=0.1 \text{ m}^3$   $p_1=10 \text{ MPa}$   $T_1=293\text{K}$   $T_2$   $p_2=5 \text{ MPa}$  kg K J

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The required text is: Introduction to Physical Gas Dynamics W. G. Vincenti and C. H. Kruger Course Structure The lectures will emphasize concepts and theory and will follow the resource material fairly closely. In some cases, the textbook will be supplemented by additional notes and handouts.

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